**Can You Tell the Difference?**

Ionic and Covalent Compounds

Part 1: Formulas🡪 Names

Name the following compounds based on their chemical formulas. First you must determine if they are ionic compounds or covalent compounds before you can name them accordingly.

|  |  |  |
| --- | --- | --- |
| **Ionic or Covalent:** | **Formula:** | **Name:** |
|  | CaO |  |
|  | Na2SO4 |  |
|  | P2O5 |  |
|  | CO |  |
|  | NO2 |  |
|  | PbCl4 |  |
|  | SO2 |  |
|  | MgSO3 |  |
|  | C3H8 |  |
|  | Li2S |  |
|  | NH4Cl |  |
|  | PF3 |  |
|  | Al2O3 |  |
|  | Pt(NO3)2 |  |
|  | SiCl4 |  |

Part 2: Names🡪 Formulas

Write the formulas for the following compounds based on their names. First you must determine if they are ionic compounds or covalent compounds before you can formulate them accordingly.

|  |  |  |
| --- | --- | --- |
| **Ionic or Covalent:** | **Formula:** | **Name:** |
|  |  | Chromium III Bromide |
|  |  | Dinitrogen trioxide |
|  |  | Carbon dioxide |
|  |  | Silver acetate |
|  |  | Trisulfur dihydride |
|  |  | Tetracarbon decafluoride |
|  |  | Scandium I hydroxide |
|  |  | Copper II chloride |
|  |  | Sulfur octobromide |
|  |  | Lithium carbonate |
|  |  | Iron IV cyanide |
|  |  | Dibromine difluoride |
|  |  | Nitrogen trichloride |
|  |  | Copper III sulfite |
|  |  | Iodine monofluoride |